

statement 1 the complete electrolysis of 45 grams of water will yield 40 grams of H_2 and 5 grams of O_2 statement 2 water is composed of hydrogen and oxygen in a ratio of 8 : 1 by mass

in the electrolysis of acidified water a name the gas collected at i anode and ii cathode b why is the volume of gas collected at one electrode double than that at the other electrode c what would

faraday's first law of electrolysis the amount of substance that undergoes oxidation or reduction at each electrode during electrolysis is directly proportional to the amount of electricity that passes

in electrolysis water is decomposed in the presence of electricity to its components the reaction is shown as below $2H_2O \rightarrow 2H_2 + O_2$ as you can see that water splits into 2 molecules of

a direct current deposits 54 g of silver atomic mass 108 during electrolysis how much aluminium atomic mass 27 would be deposited from aluminium chloride solution by the same amount of

click here to get an answer to your question during an electrolysis of conc H_2SO_4 perdisulphuric acid $H_2S_2O_8$ and O_2 form in equimolar

after electrolysis of a sodium chloride solution with inert electrodes for a certain period of time 600 ml of the solution was left which was found to be 1 N in NaOH during the same period 31.75 g of

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